



Chemistry Worksheet

Class XII The p-Block Elements

- Q1. Write the order of thermal stability of the hydrides of group 16 elements.
- Q2. Why H_2S is less acidic than H_2Te ?
- Q3. Give two examples to show the anomalous behavior of fluorine.
- Q4. Sea is the greatest source of some halogens. Comment.
- Q5. Name two poisonous gases which can be prepared from chlorine gas.
- Q6. Why has it been difficult to study the chemistry of radon?
- Q7. List the uses of neon and argon gases.
- Q8. Why do noble gases have comparatively large atomic sizes?
- Q9. Why are halogens coloured?
- Q10. Draw the structure of:
a. BrF_3 b. $\text{H}_2\text{S}_2\text{O}_7$
- Q11. Account the following observations.
a. SF_4 is easily hydrolysed whereas SF_6 is not easily hydrolysed.
b. Chlorine water is a powerful bleaching agent.
c. Bi(V) is a stronger oxidizing agent than Sb(V) .
- Q12. What happened when XeF_6 undergoes partial hydrolysis.
- Q13. What inspired N. Bartlett for carrying out reaction between Xe and PtF_6 ?
- Q14. Arrange the following in the order of property indicated against each set:
a. $\text{F}_2, \text{I}_2, \text{Br}_2, \text{Cl}_2$ (increasing bond dissociation enthalpy)
b. $\text{NH}_3, \text{AsH}_3, \text{SbH}_3, \text{BiH}_3, \text{PH}_3$ (decreasing base strength)
- Q15. Complete the following reactions-
a. $\text{Cl}_2 + \text{NaOH}$ (cold and dilute) \rightarrow
b. $\text{Fe}^{3+} + \text{SO}_2 + \text{H}_2\text{O} \rightarrow$